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⑦① Applicant: **Asahi Kasei Kogyo Kabushiki Kaisha**
2-6, Dojimahama 1-chome Kita-ku
Osaka-shi Osaka 530(JP)

⑦② Inventor: **Aoshima, Atsushi**
42-5, Sumiregaoka Kohoku-ku
Yokohama-shi Kanagawa-ken(JP)

⑦③ Inventor: **Yamamatsu, Setsuo**
1434-15, Atsuhara
Fuji-shi Shizuoka-ken(JP)

⑦④ Inventor: **Tonomura, Shoichiro**
A-25, Endo-Apartment 68, Aoshimacho
Fuji-shi Shizuoka-ken(JP)

⑦⑤ Representative: **Strehl, Schübel-Hopf, Schulz**
Widenmayerstrasse 17 Postfach 22 03 45
D-8000 München 22(DE)

⑥⑤ Process for purification of polyalkylene ether.

⑥⑦ A process for purifying a polyalkylene ether, which comprises mixing a polyalkylene ether or a mixture of a polyalkylene ether and an organic solvent containing a heteropoly acid and/or its salt with at least one solvent for purification selected from hydrocarbons having 3 to 15 carbon atoms or halogenated hydrocarbons having 1 to 15 carbon atoms and separating by precipitation the phase composed mainly of the heteropoly acid and/or its salt, and a process wherein the polyalkylene ether or a mixture of a polyalkylene ether and an organic solvent containing a heteropoly acid and/or its salt is brought into contact with a solid adsorbent capable of adsorbing the heteropoly acid and/or its salt in the presence of at least one solvent for purification as described above, either further after the separation according to the process described above, or in the state unseparated.

According to the present invention, the heteropoly acid and/or its salt dissolved in the polyalkylene ether or the mixture of the polyalkylene ether with the organic solvent can be removed efficiently in the form which can be reused. Further after the treatment with an adsorbent, the concentration of heteropoly acids in the polyalkylene ether can be reduced to extremely lower.

EP 0 181 621 A2

DESCRIPTION

PROCESS FOR PURIFICATION OF POLYALKYLENE ETHER

TECHNICAL FIELD

5 The present invention relates to a process for
purifying a polyalkylene ether by separating and
recovering a heteropoly acid and/or its salt from a
polyalkylene ether or a mixture of a polyalkylene ether
with an organic solvent containing a heteropoly acid
and/or its salt. The present invention provides a
10 process for purifying a polyalkylene ether by separating
and recovering heteropoly acids dissolved from a poly-
merizing solution containing as main components a
polyalkylene ether and an unreacted monomer obtained by a
polymerization with heteropoly acids as a catalyst.

15 BACKGROUND ART

Polyalkylene ethers are recently highlighted as the
starting material for Spandex, polyurethane elastomers,
polyester elastomers, etc., their amounts used are simply
increasing and various improvements have been attempted
20 in a production process and a purification process.

Among them, there is recently reported a process for
synthesis of a polyalkylene ether by polymerization of
a cyclic ether such as tetrahydrofuran (hereinafter
abbreviated as THF) with the use of a heteropoly acid as
25 the catalyst (European Pat. Appln. No. 84.105 720.1).
According to this process, after the reaction, most of
the catalyst can be easily recovered from the polymer-
izing solution by way of phase separation, and yet the
process has the advantage of being capable of using the
30 catalyst in recycle. However, since the catalyst is

0181621

partially dissolved in the polymerizing solution, in
order to obtain a high purity polymer and lower the
production cost of a polymer, it is required to perform
the operations for separation and recovery of the
5 catalyst dissolved in the polymerizing solution.

Ordinarily, for removal of the acid catalyst in the
polymer, there is employed a method such as extraction
with water, or filtration or washing out the salt formed
with addition of a basic substance. However, there is a
10 strong interaction between a polyalkylene ether and a
heteropoly acid, and hence an extraction with water
cannot remove the catalyst efficiently but requires a
large amount of water. Further, a polyalkylene ether
will readily form an emulsion with water, whereby the
15 problem is caused that phase separation can be effected
with difficulty. Also, when a neutralization reaction is
carried out, an expensive heteropoly acid becomes no
longer available for reuse to result disadvantageously in
increased cost for a production of the polymer.

20 Further, even when washing with water or a neutrali-
zation with a basic substance may be performed, it is
very difficult to reduce the heteropoly acid content to 0.1 wt %
or less, particularly 0.01 wt. % or less based on a
polyalkylene ether.

25 In the light of the state of the art as described
above, the present inventors have made intensive studies
about a process for removing efficiently a heteropoly
acid in a polyalkylene ether, and consequently found a
phenomenon that the heteropoly acid is precipitated (in case
30 of a liquid mixture of the polyalkylene ether and the hetero-
poly acid), or the heteropoly acid alone remains as a solid
phase (in case of a solid mixture of the polyalkylene ether
and the heteropoly acid) when a specific solvent for puri-
fication is added to the polyalkylene ether containing the
35 stated impurity.

0181621

And, further, various methods for effectively removing the heteropoly acid remaining minutely in a polyalkylene ether even after the above treatment have also been investigated. As a consequence, it has been found that, by treatment with the use of an adsorbent in co-presence of a specific solvent for purification, heteropoly acids can be removed to a very low concentration. The present invention has been accomplished on the basis of such findings.

10 DISCLOSURE

The present invention provides a process for purifying a polyalkylene ether, which comprises mixing a polyalkylene ether or a mixture of an organic solvent and a polyalkylene ether containing a heteropoly acid and/or its salt (hereinafter sometimes abbreviated as "heteropoly acids") with at least one solvent for purification selected from hydrocarbons having 3 to 15 carbon atoms or halogenated hydrocarbons having 1 to 15 carbon atoms (hereinafter sometimes abbreviated as "a solvent for purification") and separating the phase composed mainly of the heteropoly acid and/or its salt, and, if desired, thereafter bringing the polyalkylene ether or the mixture of an organic solvent and a polyalkylene ether containing heteropoly acids into contact with a solid adsorbent capable of adsorbing the heteropolyacids in the presence of the above-described solvent for purification.

The heteropoly acid and its salt in the present invention refers comprehensively to oxy acids formed by condensation of at least one oxide of Mo, W and V with oxy acids of other elements such as P, Si, As, Ge, B, Ti, Ce, Co, etc. and salts of the oxy acids, the atomic ratio of the former to the latter being 2.5 to 12.

0181621

Examples of these heteropoly acids and its salts may include 12-molybdophosphoric acid, 5-molybdo-2-phosphoric acid, 12-tungstophosphoric acid, 12-molybdotungstophosphoric acid, 6-molybdo-6-tungstophosphoric acid, 5 12-molybdovanadophosphoric acid, 11-molybdo-1-vanadophosphoric acid, 12-molybdotungstovanadophosphoric acid, 12-tungstovanadophosphoric acid, 12-molybdoniobophosphoric acid, 12-tungstosilicic acid, 12-molybdosilicic acid, 12-molybdotungstosilicic acid, 12-molybdotungstovanadosilicic acid, 12-tungstoboric acid, 12-molybdoboric acid, 10 12-molybdotungstoboric acid, 12-molybdovanadoboric acid, 12-molybdotungstovanadoboric acid, 9-molybdonickelic acid, 6-molybdocobaltic acid, 6-tungstocobaltic acid, 11-molybdoarsenic acid, 12-tungstoarsenic acid, 15 12-tungstogermanic acid, 18-tungsto-2-arsenic acid, 18-molybdo-2-phosphoric acid, 9-molybdophosphoric acid, 18-tungsto-2-phosphoric acid, 12-titanomolybdic acid, 12-ceriomolybdic acid, 18-molybdo-2-phosphoric acid and salts thereof.

20 Among them, preferable examples include 12-molybdo-phosphoric acid, 18-molybdo-2-phosphoric acid, 9-molybdophosphoric acid, 12-tungstophosphoric acid, 18-tungsto-2-phosphoric acid, 11-molybdo-1-vanadophosphoric acid, 12-molybdotungstophosphoric acid, 25 6-molybdo-6-tungstophosphoric acid, 12-molybdotungstovanadophosphoric acid, 12-tungstovanadophosphoric acid, 12-molybdosilicic acid, 12-tungstosilicic acid, 12-molybdotungstosilicic acid, 12-molybdotungstovanadosilicic acid, 12-tungstoboric acid, 12-molybdoboric acid, 30 12-molybdotungstoboric acid, 12-molybdovanadoboric acid, 12-molybdotungstovanadoboric acid, 12-tungstogermanic acid and 12-tungstoarsenic acid.

The kinds of the salts are not particularly limited. For example, there may be employed metal salts of the

0181621

metals belonging to the group I of the periodic table such as Li, Na, K, Rb, Cs, Cu, Ag, Au, etc., of the group II such as Mg, Ca, Sr, Ba, Zn, Cd, Hg etc., of the group III such as Sc, La, Ce, Al, Ga, In, etc., of the group VIII such as Fe, Co, Ni, Ru, Pd, Pt, etc. and other metals such as Sn, Pb, Mn, Bi, etc., or ammonium salts, amine salts and others.

Typical examples of these salts may include

- 12-tungstophosphate-1-lithium ($\text{LiH}_2\text{PW}_{12}\text{O}_{40}$), 12-tungsto-
- 10 phosphate-2-lithium ($\text{Li}_2\text{HPW}_{12}\text{O}_{40}$), 12-tungsto-
phosphate-1-sodium ($\text{NaH}_2\text{PW}_{12}\text{O}_{40}$), 12-tungstophosphate-
2-sodium ($\text{Na}_2\text{HPW}_{12}\text{O}_{40}$), 12-tungstophosphate-2-potassium
($\text{K}_2\text{HPW}_{12}\text{O}_{40}$), 12-tungstophosphate-2-cesium ($\text{Cs}_2\text{HPW}_{12}\text{O}_{40}$),
12-tungstopnosphate-2-silver ($\text{Ag}_2\text{HPW}_{12}\text{O}_{40}$), 12-tungsto-
- 15 phosphate-1-magnesium ($\text{MgHPW}_{12}\text{O}_{40}$), 12-tungsto-
phosphate-1-calcium ($\text{CaHPW}_{12}\text{O}_{40}$), 12-tungstophosphate-
1-zinc ($\text{ZnHPW}_{12}\text{O}_{40}$), 12-tungstophosphate-1-nickel
($\text{NiHPW}_{12}\text{O}_{40}$), 12-tungstosilicate-1-nickel ($\text{NiH}_2\text{SiW}_{12}\text{O}_{40}$),
12-tungstosilicate-2-lithium ($\text{Li}_2\text{H}_2\text{SiW}_{12}\text{O}_{40}$),
- 20 12-tungstosilicate-2-silver ($\text{Ag}_2\text{H}_2\text{SiW}_{12}\text{O}_{40}$), 12-tungsto-
silicate-1-magnesium ($\text{MgH}_2\text{SiW}_{12}\text{O}_{40}$), 12-tungsto-
silicate-1-aluminum ($\text{AlHSiW}_{12}\text{O}_{40}$), 12-tungstosilicate-
1-indium ($\text{InHSiW}_{12}\text{O}_{40}$), 12-tungstosilicate-1-gallium
($\text{GaHSiW}_{12}\text{O}_{40}$), 12-molybdophosphate-1-lithium
- 25 ($\text{LiH}_2\text{PMo}_{12}\text{O}_{40}$), 12-molybdophosphate-1-magnesium
($\text{MgHPMo}_{12}\text{O}_{40}$), 12-tungstophosphate-2-ammonium
($(\text{NH}_4)_2\text{HPW}_{12}\text{O}_{40}$), 12-tungstosilicate-1-tetramethylamine
($\text{N}(\text{CH}_3)_4\text{H}_3\text{SiW}_{12}\text{O}_{40}$), 12-tungstophosphate-1-iron
($\text{FePW}_{12}\text{O}_{40}$), 12-tungstophosphate-1-bismuth ($\text{BiPW}_{12}\text{O}_{40}$),
- 30 12-tungstophosphate-1-aluminum ($\text{AlPW}_{12}\text{O}_{40}$), 12-tungsto-
phosphate-1-chromium ($\text{CrPW}_{12}\text{O}_{40}$), 12-tungstophosphate-
1-gallium ($\text{GaPW}_{12}\text{O}_{40}$), 12-tungstophosphate-1-indium
($\text{InPW}_{12}\text{O}_{40}$) and the like.

0181621

The heteropoly acid and/or its salt may be either a single compound or a mixture of two or more kinds. The amount of the heteropoly acids to be dissolved may preferably 0.001 % by weight to 50 % by weight, more preferably 0.001 % by weight to 10 % by weight, based on the polyalkylene ether or the mixture of the polyalkylene ether and the organic solvent.

The polyalkylene ether in the present invention is a polymer having polyoxyalkylene chains of the formula:

10
$$-[O - (CH_2)_n]_m-$$
 (wherein $n \geq 1$, $m \geq 2$)
or polyoxyalkylene chains having substituents such as alkyl group, halogenated alkyl group, etc. in the molecule, which is not particularly limited. Its molecular weight is not particularly limited, but preferably within the range from 500 to 5000. Such polymers may include polymers of THF, copolymers of THF with other cyclic ethers, copolymers of THF with polyhydric alcohols, and the like. Further, copolymers of THF with other cyclic ethers and polyhydric alcohols may also be
15 included.
20

The organic solvent mixed with the polyalkylene ether is not particularly limited, but it is possible to exemplify cyclic ethers, polyhydric alcohols which are the monomers constituting the polyalkylene ether or
25 alcohols generally employed as polymerization terminators. Also, water may be contained in a small amount.

The polymerization mixture obtained by the reaction of THF alone or THF with a cyclic ether, THF with a
30 polyhydric alcohol or THF with a cyclic ether and a polyhydric alcohol with the use of heteropoly acids as polymerization catalysts is a suitable example, and the

fact that the polymerization as described above proceeds
is disclosed in the above-identified patent application
and in the European Pat. Appln. No. 85.103 690.5.

The polymerization mixture obtained by the above
5 polymerization reaction is frequently separated into two
phases comprising the phase composed mainly of the
polymer and the unreacted monomer and the phase composed
mainly of the heteropoly acid catalyst and, in this case,
the present process may be applied to the polymerization
10 mixture from which the catalyst phase has been previously
removed. The reaction mixture in the above applications
contains generally 0.1 to 3 % by weight of heteropoly
acids dissolved therein.

Examples of hydrocarbons having 3 to 15 carbon atoms
15 or halogenated hydrocarbons having 1 to 15 carbon atoms
to be used for the precipitation of heteropoly acids
("solvents for purification") are cyclopentane, cyclo-
hexane, cycloheptane, cyclooctane, methylcyclo-pentane,
methylcyclohexane, chlorocyclohexane, propane, butane,
20 pentane, hexane, heptane, octane, nonane, decane,
undecane, dodecane, chloroform, trichlorotrifluoroethane,
trichlorofluoromethane, benzene, toluene, chlorobenzene,
and the like. Propane, butane, etc. are required to be
handled under pressurization. Particularly preferred are
25 pentane, hexane, heptane, octane, nonane, decane,
undecane, dodecane, cyclohexane, cyclopentane, cyclo-
octane, cyclononane and cyclodecane, since they have a higher
performance of precipitating heteropoly acids effec-
tively even in small amounts.

30 When the present process is to be applied to a
reaction mixture comprising a polyalkylene ether and
unreacted THF, for recovery of the unreacted THF from a
solvent for purification added, the solvent for

0181621

purification should preferably have a boiling point higher than THF and not form an azeotropic mixture therewith. Although n-hexane has been known to form an azeotropic mixture with THF, n-heptane, octane and nonane are found in the present invention to form no azeotropic mixture with THF. Further, since these solvents can be rather easily recovered from the mixture with polyalkylene ether, n-heptane, octane and nonane are particularly preferred as the solvent. A solvent having 16 or more carbon atoms is not preferred, because it has a high boiling point and can be recovered by separation from a polyalkylene ether with difficulty. Even when no azeotropic mixture will be formed, a solvent having not more than 7 carbon atoms requires a large number of necessary stages of a rectifying tower, and, although depending on the amount of the hydrocarbon employed, the liquid within or at the bottom of the rectifying tower sometimes separates into two phases, that is, the polyalkylene ether phase and the hydrocarbon phase. When the separation into two phases occurs within the rectifying tower, the stage efficiency is decreased and also undesirable foaming will readily occur. When using octane or nonane, two phase separation within the rectifying tower will hardly occur and therefore octane and nonane are most preferred as the solvent.

The details about the function of precipitating heteropoly acids are not clear, but it may be estimated that low solubility of heteropoly acids in the above-described solvent for purification and an appropriate compatibility of the solvent for purification with a polyalkylene ether are each one of causes therefor.

The amount of the solvent for purification to be used for precipitation of heteropoly acids depends on the coexisting amounts of organic solvents such as unreacted

0181621

monomers, but may generally be at least equal to, preferably two-fold or more of the weight of organic solvents such as unreacted monomers. Also, it is added in an amount generally of 0.5-fold or more, preferably
5 1.0-fold or more, relative to the polyalkylene ether. When the content of the organic solvents such as unreacted monomers is small, the solvent for purification added may sometimes cause a phase separation from the polyalkylene ether phase, but the effect is small even if
10 the solvent for purification may be used in an amount over the amount which will cause a phase separation.

When organic solvents such as unreacted monomers coexist in large amounts, a large amount of the solvent for purification is required to be used for precipitation
15 of heteropoly acids. For this reason, it is preferable to recover at least a portion of the organic solvent by concentrating the mixture to a content of organic solvents of 90 wt. % or less, more preferably 50 wt. % or less, before adding the solvent for purification to precipitate
20 heteropoly acids.

For mixing of the solvent for purification, a conventional stirring means, a static mixer, etc. may be used. After mixing, the mixture may be left to stand, whereby the phase consisting mainly of heteropoly acid or
25 its salt is precipitated to form two phases together with the polymer solution phase. In the phase consisting mainly of heteropoly acids as the lower phase, a small amount of polymer and organic solvents such as unreacted monomers and a very minute amount of the solvent for
30 purification are contained, although the amounts may differ more or less depending on the practicing conditions.

Mixing and standing may be carried out at a temperature at which the operation can easily be done within the range which will not coagulate the polymer. If the temperature is too high, the solubility of the heteropoly acid is increased and therefore no higher temperature than necessary should be used. Generally, the operations are practiced at 25 to 60 °C. The mixture may be left standing for about 0.1 hour to 100 hours. The standing time can be shortened by use of a coalescer, a liquid cyclone, a centrifugal machine, etc. Also, by filtration by use of a membrane filter, the catalyst can be separated. On completion of a phase separation, the lower phase containing heteropoly acids at a high concentration is separated from the upper phase. By removing the organic solvents and the solvent for purification by an appropriate method such as distillation, polymers having decreased concentration of heteropoly acids can be obtained. The lower phase can be reused as such as the polymerization catalyst.

When the upper phase consists mainly of a polyalkylene ether, THF and a hydrocarbon solvent for purification which has a higher boiling point than THF and does not form an azeotropic mixture therewith, although somewhat different depending on the amount of the solvent for purification, only THF can be recovered from the rectifying tower and the remaining mixture can be separated by cooling at an appropriate temperature into two phases of the phase consisting mainly of the polyalkylene ether and the phase consisting mainly of the solvent for purification, whereby most of the solvent for purification can be recovered by phase separation. The solvent recovered can be used as such, or after a purification operation such as distillation, again as the solvent for separating the phase composed mainly of the heteropoly acid and/or its salt.

0181621

By practicing the above-described process of the present invention, heteropoly acids can be removed and recovered efficiently in a form which can be reused, and the heteropoly acid concentration in the polyalkylene ether can be reduced to 0.5 wt. % or less, even 0.01 wt. % or less under optimum practicing conditions.

Polyalkylene ether having reduced or low concentration of heteropoly acids of 0.5 wt. % or less, respectively with or without practice of the above-described operation, may further be purified to lower the heteropoly acid concentration, by the step of contacting with a solid adsorbent in the co-presence of at least one solvent selected from hydrocarbons having 3 to 15 carbon atoms or halogenated hydrocarbons having 1 to 15 carbon atoms ("a solvent for purification"). The solvent for purification to be co-present may be the same as that used for separating the phase composed mainly of heteropoly acids. Usually, the solvent for purification is added to a polymerizing solution comprising the polyalkylene ether, unreacted monomers and heteropoly acids to separate the phase consisting mainly of heteropoly acids, followed by contact with the solid adsorbent. When the concentration of the dissolved heteropoly acid is low or when the amount of the solvent for purification added is small, the phase consisting of heteropoly acids may sometimes not be separated and, in this case, only the operation of contacting the polymerizing solution with the solid adsorbent is practiced.

The amount of the hydrocarbon or the halogenated hydrocarbon ("a solvent for purification") to be co-present when contacting with the solid adsorbent may be preferably 0.05-fold or more, more preferably 0.5-fold or more of the polyalkylene ether. When an organic solvent such as THF is admixed, it is preferable to add a

solvent for purification in an amount at least equal to the amount of the organic solvent. When a polyalkylene ether alone or a small amount of the organic solvent is mixed, the solvent for purification may be phase
5 separated from the polyalkylene ether phase, but the effect is small even if the solvent for purification may be used in an amount to such an extent to cause the phase separation. On the other hand, when the solvent for purification is previously admixed in a suitable amount, it
10 is not particularly required to be added, but it may also be added, if desired.

The kind of the solid adsorbent is not particularly limited, provided that it can adsorb the heteropoly acid and/or its salt, but preferred examples are one or a
15 mixture of two or more selected from (A) activated charcoal, (B) aluminum oxides such as alumina, silica-alumina, etc., (C) oxides, hydroxides or carbonates of magnesium, calcium and rare earth metals and (D) basic ion-exchange resins.

20 The activated charcoal used is not particularly limited but may be any of coconut shell type, coal type, petroleum type, charcoal type. Also, as for a basic ion-exchange resin, it is not particularly limited but may preferably be of microporous type.

25 Among them, activated charcoal, alumina and silica-alumina have the advantages of high adsorption ability and being hardly crushable in the case of treatment when the adsorbent is packed in a column. Among solid basic compounds as described above in (C), CaO , Ca(OH)_2 and
30 CaCO_3 are preferably cheap, but MgO has the advantage of low solubility. A basic ion-exchange resin is lower in adsorption ability as compared with other adsorbents, is required to be used in a large amount and therefore has

0181621

the drawback that the impurities from the resin may be dissolved out.

5 The amount of the adsorbent to be used, which may differ depending on the amount of heteropoly acids dissolved, may be 2 to 5000-fold weight, more preferably 10 to 1000-fold weight of the amount of heteropoly acids dissolved. As the amount used is larger, the concentration of heteropoly acids after treatment will be lowered.

10 The treatment temperature is not particularly limited, but a temperature for an easy operation within the temperature range in which the liquid to be treated has an appropriate viscosity is selected. For example, for polyoxytetramethylene glycol or a copolymer of
15 polyalkylene ether glycol of THF with other cyclic ethers or THF with diols, the temperature may be 20 to 150 °C, more preferably 30 to 100 °C.

As the treatment method, there may be employed any of the conventional methods such as the method in which
20 after mixing and stirring the liquid to be treated with adsorbent powder or particles, the adsorbent is separated by filtration, and the method in which the liquid to be treated is passed through a column packed with the adsorbent, and the method may be either batchwise or
25 continuous. The treatment time, which is not particularly limited, may generally be 0.05 to 10 hours, particularly 0.1 to 3 hours. The adsorbent used can be utilized, as long as it has the adsorption capacity, as such or after a regeneration treatment.

30 After the adsorption treatment, the organic solvent and the solvent for purification co-existing can be evaporated by a distillation at normal or reduced

0181621

pressure to provide a polyalkylene ether lowered in a concentration of heteropoly acids.

According to the present invention, the concentration of heteropolyacids can be lowered to 0.0001 wt. % or less. Besides, since a dissolution of impurities out of the adsorbent can also be suppressed, by removing the solvent by a distillation, etc. after the adsorption treatment, a polyalkylene ether extremely small in impurities can be obtained.

10 Such effects may be considered to be attributable by the co-operating action of the above solvent for purification and the adsorbent selected. More specifically, it may be estimated that, through the presence of the solvent for purification, the affinity of the heteropoly
15 acid for a polyalkylene ether is lowered to strengthen the adsorbing action of the adsorbent for the heteropoly acid or its salt, simultaneously with a suppression of a dissolution of the components out of the adsorbent
20 itself. The solvent for purification may also be considered to promote a diffusion of the heteropoly acid toward the surface of adsorbent.

By a practice of the present invention, as already mentioned above, the heteropoly acid and/or its salt dissolved in the polyalkylene ether or the mixture of the
25 polyalkylene ether with the organic solvent can be removed efficiently in the form which can be reused. Further after the treatment with an adsorbent, the concentration of heteropoly acids in the polyalkylene ether can be reduced to 0.0001 wt. % or lower.

30 After applying the treatment with the adsorbent according to the present process, unreacted monomers such as THF can be recovered by a distillation, whereby

0181621

scale formation within the distillation tower can be suppressed.

Further, the polyalkylene glycol treated according to the present process suffers little from deterioration even when heat is applied and therefore can be distilled by means of a thin film type molecular distillation apparatus at 100 to 180 °C, more preferably at 120 to 160 °C, at a reduced pressure preferably of 3 Torr or lower, whereby the organic solvent and the solvent for purification, as a matter of course, and further the linear chain or cyclic oligomer components contained in the polyalkylene glycol can be removed.

0181621

Example 1

Into a 5.0-liter vessel equipped with a stirring means and a reflux condenser were charged 2.0 kg of THF containing 300 ppm of water and 1.0 kg of 12-tungsto-
5 phosphoric acid ($H_3PW_{12}O_{40} \cdot 3H_2O$) controlled to the number of crystal water of 3 by heating in an electric furnace at 150 °C for 2 hours. The reaction system thereby
10 formes two liquid phases, that is, the catalyst liquid phase and the THF phase. Thereafter the stirring was continued at a temperature of 60 °C for 4 hours for a
polymerization, the mixture was left to stand to be separated into two liquid phases, and the catalyst phase
was recovered as the lower phase. The composition of the THF phase as the upper phase was analyzed to find that 23
15 % by weight of polyoxytetra-methylene glycol (hereinafter abbreviated as "PTMG") and about 2.8 % by weight of 12-tungstophosphoric acid were present.
Next, a part of the unreacted THF was recovered to obtain 820 g of a THF solution containing 50 % by weight of
20 PTMG. The solution was found to contain 6 % by weight of 12-tungstophosphoric acid dissolved therein. To this solution was added 1.0 kg of n-octane, and the mixture after being stirred was left to stand at 30 °C for 30
hours, whereafter the catalyst liquid phase was pre-
25 cipitated to be separated into two layers, with the upper phase becoming colorless and transparent. The weight of the catalyst phase recovered from the lower phase was 91 g, containing about 50 % by weight of 12-tungsto-
phosphoric acid, with the balance being PTMG, THF and a
30 very minute amount of n-octane. Subsequently, only THF was recovered by a distillation from the upper phase solution. Within the distillation tower and at the distillation tower bottom, the liquid was not separated into two phases, and no foaming was observed within the
35 distillation tower. The bottom liquid in the

0181621

distillation tower was withdrawn and left to stand at 40
°C, whereby n-octane was phase separated to be recovered
in an amount of about 0.8 kg. Then, by removing the
n-octane dissolved in PTMG by a distillation at a reduced
5 pressure, PTMG in which the 12-tungstophosphoric acid is
reduced to 50 ppm was obtained. The n-octane recovered
by a phase separation was used as such and after a
distillation as the solvent for precipitation of the
heteropoly acid. As the results, in both cases, the
10 solvents had sufficient function of precipitating the
heteropoly acid. Further, when the catalyst phase
recovered is recycled as the catalyst for polymerization
of THF, a polymerization of THF was found to proceed.

In this Example and in the following Examples, a
15 microanalysis of the heteropoly acid or its salt was
conducted by analysis of W, Mo, etc. according to the
fluorescent X-ray analytical method. It is not clear
whether W or Mo remaining minutely in the polyalkylene
ether after the adsorption treatment still remains the
20 structure of the original heteropoly acid.

Comparative Example 1

When THF was evaporated by reduced pressure distil-
lation from the THF solution containing 6 % by weight of
12-tungsto-phosphoric acid and 50 % by weight of PTMG
25 obtained in the same manner as Example 1, no precipita-
tion of the catalyst phase occurred at all, but the
existing 12-tungstophosphoric acid was all dissolved in
PTMG to give PTMG containing 12 % by weight of
12-tungsto-phosphoric acid.

30 Examples 2 - 14

Various solvents as indicted in Table 1 were added

0181621

to the polymers containing the heteropoly acid or the salt thereof or mixtures containing the said polymers and THF as indicated in Table 1, followed by stirring, and they were left to stand at 30 °C for 50 hours. The catalyst phase sedimented was separated and recovered, and the upper phase was desolventized to provide a polymer reduced in a concentration of the heteropoly acid or its salt. The results are shown in Table 1. As for the n-butane-added system of Example 5, the operation was conducted under a liquefying condition by using a pressure vessel. The PTMG employed in Example 6 contained about 1.0 % by weight of water.

Table 1

Example	Composition of liquid to be treated			Solvent added		Meteropoly acid conc. in the purified polymer(ppm)
	Polymer	Meteropoly acid		THF	Weight (g)	
	Kind	Weight(g)	Kind	Weight(g)	Kind	
2		200	H ₃ PMo ₁₂ O ₄₀	15	150 iso-octane	80
3		200	H ₄ SiW ₁₂ O ₄₀	10	150 cyclohexane	200
4	PTMG	200	H ₄ SiMo ₁₂ O ₄₀	15	150 neo-pentane	120
5		200	H ₃ PW ₁₂ O ₄₀	20	200 n-butane	200
6		200	H ₃ PW ₁₂ O ₄₀	20	0 n-octane	700
7		200	H ₃ PW ₁₂ O ₄₀	20	0 n-heptane	800
8	Ethylene glycol - THF copolymerized polyether glycol (THF content 95 wt. %)	100	H ₄ SiW ₁₂ O ₄₀	10	100 n-undecane	80
9		100	H ₃ PW ₁₂ O ₄₀	12	100 toluene	350
10	Neopentyl glycol - THF copolymerized polyether glycol (THF content 90 wt. %)	100	H ₃ PMo ₁₂ O ₄₀	8	100 n-hexane	110
11		100	H ₃ PW ₁₂ O ₄₀	18	100 benzene	380
12	Propylene oxide - THF copolymerized polyether glycol (THF content 80 wt. %)	100	H ₄ SiW ₁₂ O ₄₀	11	150 Freon-113	200
13		100	H ₃ PW ₁₂ O ₄₀	13	150 chloroform	600
14		100	NaH ₂ PW ₁₂ O ₄₀	10	100 n-octane	500

0181621

0181621

Example 15

To a mixture comprising about 390 g of PTMG, about 200 g of THF, about 1.0 kg of n-octane and 12-tungstophosphoric acid minutely dissolved therein obtained after
5 the catalyst separation following the operation as described in Example 1, 1.0 g of MgO powder was added, followed by stirring at 40 °C for one hour. Then, after MgO powder was removed with a Teflon membrane filter, most of THF and n-octane were recovered by a reduced
10 pressure distillation. As the next step, by means of a thin film system molecular distillation device, a distillation was carried out at 140 °C, 0.1 Torr to evaporate the solvent, the water and the chain and cyclic oligomer components contained in polyoxytetramethylene
15 glycol, whereby a purified PTMG reduced in W content to 0.5 ppm was obtained. THF and n-octane recovered by a distillation were separated in a rectifying tower, and THF was reused as the starting material, while n-octane as the solvent for purification.

20 Comparative example 2

When a mixture comprising 390 g of PTMG, 200 g of THF and 50 ppm of 12-tungstophosphoric acid was treated in the same manner as described in Example 15, a PTMG with a W content of 20 ppm was obtained. Further, PTMG
25 was partially depolymerized and lowered in recovery.

Examples 16 - 27

Into the compositions as indicated in Table 2, comprising 100 g of a polyalkylene ether and hydrocarbon solvents containing minute amounts of heteropoly acids,
30 predetermined amounts of adsorbents as indicated in Table 2 were added, followed by stirring at 80 °C for one hour.

0181621

Then, the adsorbent was recovered by filtration, and thereafter the solvent was evaporated by a reduced pressure distillation to provide a polyalkylene ether reduced in W content. The results are shown in Table 2.

Table 2

Composition of liquid to be treated by an adsorbent					Adsorbent	W or Mo content in the purified polyalkylene ether (ppm)
Example	Polyalkylene glycol (100 g)	Heteropoly-acid fication		Solvent for purification		
		Kind	Content (ppm)		Kind	Weight (g)
16		H ₃ PMo ₁₂ O ₄₀	100	iso-octane	30 Activated charcoal (Calgon CAL)	5 1.8
17		H ₄ SiW ₁₂ O ₄₀	150	cyclohexane	50 Activated charcoal (Kreha beads BAC-MQ)	5 1.2
18	Polyoxytetramethylene glycol	H ₃ PW ₁₂ O ₄₀	200	neo-pentane	30 CaO powder	1 0.5
19		H ₃ PW ₁₂ O ₄₀	150	n-octane	30 Al ₂ O ₃ powder	0.5 0.4
20		H ₃ PW ₁₂ O ₄₀	150	n-undecane	30 Al ₂ O ₃ -SiO ₂ powder	0.5 0.6
21		H ₄ SiMo ₁₂ O ₄₀	150	toluene	80 MgO powder	1 2.5
22	Ethylene glycol - THF copolymerized polyether glycol (THF content 95 wt. %)	H ₃ PMo ₁₂ O ₄₀	180	n-hexane	Activated charcoal (Takeda-Shirasagi A powder)	1 0.2
23		H ₃ PW ₁₂ O ₄₀	180	Freon-113	50 Al ₂ O ₃ powder	1 1.2
24	Neopentyl glycol - THF copolymerized polyether glycol (THF content 90 wt. %)	H ₃ PW ₁₂ O ₄₀	150	chloroform	100 Y-zeolite powder	1 2.5
25		H ₄ SiW ₁₂ O ₄₀	150	n-heptane	30 Active clay	1 3.5
26	Propylene oxide - THF copolymerized polyether glycol (THF content 80 wt. %)	H ₄ SiW ₁₂ O ₄₀	160	n-octane	30 CeO ₂ powder	1 1.0
27		H ₃ PW ₁₂ O ₄₀	170	n-octane	30 Basic ion-exchange resin Amberlite IRA-900	15 5.5

0181621

Example 28

Into a vessel equipped with a stirring means and a reflux condenser were charged 2.0 kg of THF and 85 g of ethylene glycol. Then, 1.0 kg of 12-tungstosilicic acid
5 ($H_4SiW_{12}O_{40}$) heated at 320 °C for 3 hours to be made anhydrous was added thereto. By setting the temperature at 60 °C, stirring was continued for 8 hours and the mixture was left to stand at room temperature to be separated into two phases of the catalyst phase and the
10 THF phase. The composition of the THF phase of the upper phase was analyzed to find that it contained about 61 % by weight of THF, about 35 % by weight of a polyalkylene ether glycol in which ethylene glycol is copolymerized with THF, about 4 % by weight of 12-tungstosilicic acid
15 and about 0.3 % by weight of water. To 500 g of this mixture was added 625 g of n-octane, and the resultant mixture was stirred. The turbid mixture in which the heteropoly acid phase was separated was filtered through a Teflon membrane filter to separate the heteropoly acid
20 phase and obtain a colorless and transparent filtrate. Then, the filtrate was passed through a column of 20 cm in length packed with 100 g of granular activated charcoal (CAL produced by Calgon Co.) at a linear velocity of 20 cm/hour. Then, only THF was recovered by
25 a distillation, followed by cooling to 40 °C, whereby n-octane was phase separated to be recovered in an amount of 500 g. Subsequently, most of the n-octane dissolved in the polyalkylene ether glycol was recovered by a reduced pressure distillation, and thereafter distilled
30 by means of a thin film type molecular distillation device at 140 °C, 0.1 Torr to remove n-heptane, water, linear chain and cyclic oligomer components and other low boiling point impurities, whereby a purified polyalkylene ether glycol comprising a copolymer of THF and ethylene
35 glycol reduced in W content to 0.1 ppm was obtained. The

0181621

n-octane recovered by a phase separation was distilled and combined with the n-octane recovered by a reduced pressure distillation and utilized again as the solvent for purification.

CLAIMS

1. A process for purifying a polyalkylene ether containing at least one heteropoly acid and/or its salt as an impurity, which comprises mixing the impurity containing polyalkylene ether as such or in the form of its mixture with one or more organic solvents with at least one solvent for purification selected from hydrocarbons having 3 to 15 carbon atoms or halogenated hydrocarbons having 1 to 15 carbon atoms and separating the phase composed mainly of the heteropoly acid and/or its salt.

2. A process for purifying a polyalkylene ether containing at least one heteropoly acid and/or its salt as an impurity, which comprises contacting the impurity containing polyalkylene ether as such or in the form of its mixture with one or more organic solvents with a solid adsorbent capable of adsorbing the heteropoly acid and/or its salt in the presence of at least one solvent for purification selected from hydrocarbons having 3 to 15 carbon atoms or halogenated hydrocarbons having 1 to 15 carbon atoms, wherein the step of contacting with the solid adsorbent is optionally carried out after a separation step according to claim 1.

3. A process according to Claims 1 or 2, wherein the polyalkylene ether or the mixture of polyalkylene ether and an organic solvent is a polymerization mixture composed mainly of a polyalkylene ether having polyoxytetramethylene chains in the molecule and tetrahydrofuran, and the solvent for purification selected from hydrocarbons having 3 to 15 carbon atoms or halogenated hydrocarbons having 1 to 15 carbon atoms is at least one solvent for purification selected from hydrocarbons having not more than 15 carbon atoms which has a boiling point higher than tetrahydrofuran and does not form an azeotropic mixture therewith, tetrahydrofuran is recovered by distillation after separation of the phase composed mainly of the heteropoly acid and/or its salt,

and the residual mixture of the said polyalkylene ether and the said solvent for purification is subjected to phase separation to thereby separate the liquid composed mainly of the said polyalkylene ether from most of the
5 said solvent for purification.

4. A process according to Claim 3, wherein the polyalkylene ether having polyoxytetramethylene chains in the molecule is obtained by a polymerization with the use of the heteropoly acid and/or its salt as a catalyst.

10 5. A process according to Claim 3, wherein the solvent for purification existing in the said polyalkylene ether is further evaporated.

6. A process according to Claim 3, wherein the polyalkylene ether having polyoxytetramethylene chains in
15 the molecule is polyoxytetramethylene glycol.

7. A process according to Claims 3 or 6, wherein the polyalkylene ether or the mixture of polyalkylene ether and an organic solvent to be purified is brought into contact with a solid adsorbent capable of adsorbing
20 the heteropoly acid and/or its salt at either one of the stage after the phase composed mainly of the heteropolyacid and/or its salt is separated by precipitation by adding at least one solvent for purification selected from hydrocarbons having not more than 15 carbon atoms
25 which has a boiling point higher than tetrahydrofuran and does not form an azeotropic mixture therewith and the stage after further tetrahydrofuran is recovered by distillation or at the both stages thereof.

8. A process according to Claims 2 or 7, wherein
30 the solid adsorbent is one or a mixture of two or more adsorbents selected from:

- A) activated charcoal;
- B) aluminum oxides;
- C) oxides, hydroxides and carbonates of magnesium, calcium and rare earth metals; and
- 5 D) basic ion-exchange resins.

9. A process according to Claims 3, 6 or 7, wherein, after tetrahydrofuran is recovered by distillation, the solvent for purification which is recovered by the phase separation, either as such or
10 after the purification operation such as distillation, is utilized again as the solvent for purification for separating the phase composed mainly of the heteropoly acid and/or its salt.

10. A process according to Claims 1, 2, 3, 6 or 7,
15 wherein the separated phase composed mainly of the heteropoly acid and/or its salt is reused as the polymerization catalyst.

11. A process according to Claims 1 or 2, wherein the polyalkylene ether or the mixture of the polyalkylene
20 ether and the organic solvent containing heteropoly acid and/or its salt is the reaction mixture obtained by a polymerization of tetrahydrofuran alone or a copolymerization of tetrahydrofuran and a cyclic ether, tetrahydrofuran and a polyhydric alcohol, or tetrahydrofuran,
25 a cyclic ether and a polyhydric alcohol with the use of the heteropoly acid and/or its salt as a catalyst, the said reaction mixture from which the catalyst phase is separated when the said reaction mixture consists of two phases with a catalyst phase, or the concentrated reaction mixture obtained by recovering at least a part of
30 unreacted monomer from the said reaction mixture.

12. A process according to Claim 3, 6 or 7 wherein the hydrocarbon having not more than 15 carbon atoms which has a boiling point higher than tetrahydrofuran and does not form an azeotropic mixture therewith is one or a
5 mixture of octane and nonane.

(12)

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(71) Applicant: Asahi Kasei Kogyo Kabushiki Kaisha
 2-6, Dojimahama 1-chome Kita-ku
 Osaka-shi Osaka 530(JP)

(72) Inventor: Aoshima, Atsushi
 42-5, Sumiregaoka Kohoku-ku
 Yokohama-shi Kanagawa-ken(JP)

(72) Inventor: Yamamatsu, Setsuo
 1434-15, Atsuhara
 Fuji-shi Shizuoka-ken(JP)

(72) Inventor: Tonomura, Shoichiro
 A-25, Endo-Apartment 68, Aoshimacho
 Fuji-shi Shizuoka-ken(JP)

(74) Representative: Strehl, Schübel-Hopf, Groening, Schulz
 Widenmayerstrasse 17 Postfach 22 03 45
 D-8000 München 22(DE)

(54) Process for purification of polyalkylene ether.

(57) A process for purifying a polyalkylene ether, which comprises mixing a polyalkylene ether or a mixture of a polyalkylene ether and an organic solvent containing a heteropoly acid and/or its salt with at least one solvent for purification selected from hydrocarbons having 3 to 15 carbon atoms or halogenated hydrocarbons having 1 to 15 carbon atoms and separating by precipitation the phase composed mainly of the heteropoly acid and/or its salt, and a process wherein the polyalkylene ether or a mixture of a polyalkylene ether and an organic solvent containing a heteropoly acid and/or its salt is brought into contact with a solid adsorbent capable of adsorbing the heteropoly acid and/or its salt in the presence of at least one solvent for purification as described above, either further after the separation according to the process described above, or in the state unseparated.

According to the present invention, the heteropoly acid and/or its salt dissolved in the polyalkylene ether or the mixture of the polyalkylene ether with the organic solvent can be removed efficiently in the form which can be reused. Further after the treatment with an adsorbent, the concentration of heteropoly acids in the polyalkylene ether can be reduced to extremely lower.



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EUROPEAN SEARCH REPORT

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DOCUMENTS CONSIDERED TO BE RELEVANT			
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	CLASSIFICATION OF THE APPLICATION (Int. Cl.4)
A	GB-A-1 467 970 (I.C.I) ---		C 08 G 65/30
A	GB-A-1 369 304 (MARLES-KUHLMANN-WYANDOTTE) ---		
A	DE-A-2 746 911 (HULS) ---		
D, P A	EP-A-0 158 229 (ASAHI KASEI KOGYO K.K.) ---		
D, P A	EP-A-0 126 471 (ASAHI KASEI KOGYO K.K.) -----		
			TECHNICAL FIELDS SEARCHED (Int. Cl.4)
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The present search report has been drawn up for all claims			
Place of search THE HAGUE		Date of completion of the search 19-02-1988	Examiner STIENON P.M.E.
CATEGORY OF CITED DOCUMENTS X : particularly relevant if taken alone Y : particularly relevant if combined with another document of the same category A : technological background O : non-written disclosure P : intermediate document T : theory or principle underlying the invention E : earlier patent document, but published on, or after the filing date D : document cited in the application L : document cited for other reasons ----- & : member of the same patent family, corresponding document			

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